

# Cambridge International AS & A Level

CHEMISTRY 9701/03

Paper 3 Advanced Practical Skills

For examination from 2022

SPECIMEN CONFIDENTIAL INSTRUCTIONS

This document gives details of how to prepare for and administer the practical exam.

The information in this document and the identity of any materials supplied by Cambridge International are confidential and must NOT reach candidates either directly or indirectly.

The supervisor must complete the report at the end of this document and return it with the scripts.

#### **INSTRUCTIONS**

If you have any queries regarding these confidential instructions, contact Cambridge International stating the centre number, the syllabus and component number and the nature of the query.
 email info@cambridgeinternational.org
 phone +44 1223 553554



### General information about practical exams

Centres must follow the guidance on science practical exams given in the Cambridge Handbook.

#### Safety

Supervisors must follow national and local regulations relating to safety and first aid.

Only those procedures described in the question paper should be attempted.

Supervisors must inform candidates that materials and apparatus used in the exam should be treated with caution. Suitable eye protection should be used where necessary.

The following hazard codes are used in these confidential instructions, where relevant:

C corrosive
 HH health hazard
 F flammable
 MH moderate hazard
 T acutely toxic
 O oxidising

**N** hazardous to the aquatic environment

Hazard data sheets relating to substances used in this exam should be available from your chemical supplier.

#### Before the exam

- The packets containing the question papers must **not** be opened before the exam.
- It is assumed that standard school laboratory facilities, as indicated in the *Guide to Planning Practical Science*, will be available.
- Spare materials and apparatus for the tasks set must be available for candidates, if required.

#### **During the exam**

- It must be made clear to candidates at the start of the exam that they may request spare materials and apparatus for the tasks set.
- Where specified, the supervisor must perform the experiments and record the results as instructed.
   This must be done out of sight of the candidates, using the same materials and apparatus as the candidates.
- Any assistance provided to candidates must be recorded in the supervisor's report.
- If any materials or apparatus need to be replaced, for example, in the event of breakage or loss, this must be recorded in the supervisor's report.

#### After the exam

- The supervisor must complete a report for each practical session held and each laboratory used.
- Each packet of scripts returned to Cambridge International must contain the following items:
  - the scripts of the candidates specified on the bar code label provided
  - the supervisor's results relevant to these candidates
  - the supervisor's reports relevant to these candidates
  - seating plans for each practical session, referring to each candidate by candidate number

the attendance register.

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### Specific information for this practical exam

During the exam, the supervisor (NOT the invigilator) must do all the experiments and record the results on a spare copy of the question paper, clearly labelled 'supervisor's results'.

If chemicals are prepared in more than one batch, clearly labelled supervisor's results must be provided for each batch. The candidates using each batch must be listed on the supervisor's report.

#### **Apparatus**

- $1 \times 25 \, \text{cm}^3$  measuring cylinder
- $1 \times 250 \, \text{cm}^3 \, \text{beaker}$
- $1 \times \text{plastic}$  or cardboard cup, capacity approximately 150 cm<sup>3</sup>
- $1 \times \text{thermometer } (-10 \,^{\circ}\text{C to } +110 \,^{\circ}\text{C at } 1 \,^{\circ}\text{C})$
- $1 \times \text{stop-clock}$  (to measure to an accuracy of 1 s)
- $1 \times 25 \, \text{cm}^3$  pipette
- 1 × pipette filler
- $1 \times 50 \text{ cm}^3 \text{ burette}$
- $2 \times 150 \text{ cm}^3 \text{ or } 250 \text{ cm}^3 \text{ conical flask}$
- 1 × burette stand and clamp
- 1 × funnel (for filling burette)
- 1 × white tile
- 1 × glass rod
- 1 × test-tube holder
- 1 × hard-glass test-tube
- 6 × test-tube\*
- 2 × boiling tube\*
- 1 × test-tube rack
- 2 × teat / dropping pipette
- 1 × wash bottle containing distilled water
- 1 × pen for labelling glassware
- 1 × spatula
- 1 × Bunsen burner
- 1 × heat-proof mat

paper towels

red and blue litmus papers

aluminium foil for testing nitrate/nitrite

wooden splints

the apparatus normally used in the centre for use with limewater in testing for carbon dioxide

\* Candidates are expected to rinse and re-use test-tubes where possible. Additional tubes should be available.

© Materials
Signature Materials listed in the table must be provided to each candidate.

label	per candidate	identity	notes (hazards given in this column are for the raw materials)
FA 1	40 cm <sup>3</sup>	1.0 mol dm <sup>-3</sup> hydrochloric acid	Prepare 2.0 mol dm $^{-3}$ HC $l$ using the instructions in the current syllabus. Dilute 500 cm $^3$ in each dm $^3$ of solution.
FA 2 [F]	0.20 ± 0.01 g	magnesium ribbon	Provide 0.20 ± 0.01 g freshly cleaned magnesium ribbon <b>[F]</b> in a stoppered container. All candidates should be provided with a strip from the same batch.
FA 3	150 cm³	0.100 mol dm <sup>-3</sup> hydrochloric acid	Prepare 2.0 mol dm $^{-3}$ HC $l$ using the instructions in the current syllabus. Dilute 50 cm $^3$ in each dm $^3$ of solution.
FA 4	150 cm <sup>3</sup>	0.0472 mol dm <sup>-3</sup> sodium carbonate	Dissolve 5.00 g anhydrous Na <sub>2</sub> CO <sub>3</sub> <b>[MH]</b> or 13.50 g Na <sub>2</sub> CO <sub>3</sub> •10H <sub>2</sub> O <b>[MH]</b> in each dm³ of solution.
bromophenol blue [F][МН][НН]	5 cm <sup>3</sup>	bromophenol blue indicator	See preparation instructions in the current syllabus.
FA 5 [MH]	1.0 g	hydrated diammonium iron(II) sulfate	Provide 1.0 $\pm$ 0.1 g (NH <sub>4</sub> ) <sub>2</sub> Fe(SO <sub>4</sub> ) <sub>2</sub> •6H <sub>2</sub> O <b>[MH]</b> in a stoppered container.
FA 6	10 cm <sup>3</sup>	0.2 mol dm <sup>-3</sup> hydrated diammonium iron(II) sulfate	Dissolve 78.4 g (NH <sub>4</sub> ) <sub>2</sub> Fe(SO <sub>4</sub> ) <sub>2</sub> •6H <sub>2</sub> O <b>[MH]</b> in each dm <sup>3</sup> of solution.
FA 7	10 cm <sup>3</sup>	0.2 mol dm <sup>-3</sup> sodium nitrite	Dissolve 13.8 g NaNO $_2$ <b>[O][N][T]</b> in each dm $^3$ of solution.
aqueous hydrogen peroxide	10 cm <sup>3</sup>	'10-volume' hydrogen peroxide	Dilute 100 cm <sup>3</sup> of 100-volume $H_2O_2$ <b>[C]</b> to 1 dm <sup>3</sup> of solution.

label	per candidate	identity	notes (hazards given in this column are for the raw materials)
dilute nitric acid [C]	10 cm <sup>3</sup>	$2.0~\mathrm{mol~dm^{-3}~HNO_3}$	
aqueous sodium hydroxide [C]	10 cm <sup>3</sup>	2.0 mol dm <sup>-3</sup> NaOH	
aqueous barium chloride or aqueous barium nitrate	10 cm³	$0.1~{ m moldm}^{-3}~{ m BaC}\it l_2$ or $0.1~{ m moldm}^{-3}~{ m Ba(NO}_3)_2$	See preparation instructions in the current syllabus. If necessary, each of these reagents can be provided as a
limewater <b>[MH]</b>	10 cm <sup>3</sup>	saturated aqueous calcium hydroxide, Ca(OH) <sub>2</sub>	Invigilators must be alert to the risk of contamination and the opportunity for malpractice when using a communal supply.
aqueous acidified potassium manganate(VII) [MH]	10 cm <sup>3</sup>	$0.01 \text{ mol dm}^{-3} \text{ KMnO}_4 \text{ in}$ $0.5 \text{ mol dm}^{-3} \text{ H}_2 \text{SO}_4$	

An excess of at least 10% of each material must be prepared to cover accidental loss.

All solutions must be thoroughly mixed.

If you are unable to source any of these chemicals, you must contact Cambridge International as far as possible in advance of the exam for advice.

Materials must be labelled only as specified in the 'label' column. The identities of chemicals labelled with letter codes, e.g. FA 1, may be different from their descriptions in the question paper. Candidates must use the descriptions given in the question paper.

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# Supervisor's report

Syllabus and component number			1		
Centre number					
Centre name	 	 		 	
Time of the practical session	 	 		 	
Laboratory name/number	 	 		 	

Give details of any difficulties experienced by the centre or by candidates (include the relevant candidate names and candidate numbers).

You must include:

- any difficulties experienced by the centre in the preparation of materials
- any difficulties experienced by candidates, e.g. due to faulty materials or apparatus
- any specific assistance given to candidates.

If chemicals have been prepared in	more than one batch	, list the candidates	using each batch.
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## **Declaration**

1	Each packet that I am returning to Cambridge International contains the following items:
	the scripts of the candidates specified on the bar code label provided
	the supervisor's results relevant to these candidates
	the supervisor's reports relevant to these candidates
	seating plans for each practical session, referring to each candidate by candidate number
	the attendance register.
2	Where the practical exam has taken place in more than one practical session, I have clearly labelled the supervisor's results, supervisor's reports and seating plans with the time and laboratory name/number for that practical session.
3	I have included details of difficulties relating to this practical session experienced by the centre or by candidates.
4	I have reported any other adverse circumstances affecting candidates, e.g. illness, bereavement or temporary injury, directly to Cambridge International on a <i>special consideration form</i> .
Sigr	ned (supervisor)
Nan	me (in block capitals)
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